

Novel Method of Quantifying Radioactive Cesium-Rich Microparticles (CsMPs) in the Environment from the Fukushima Daiichi Nuclear Power Plant

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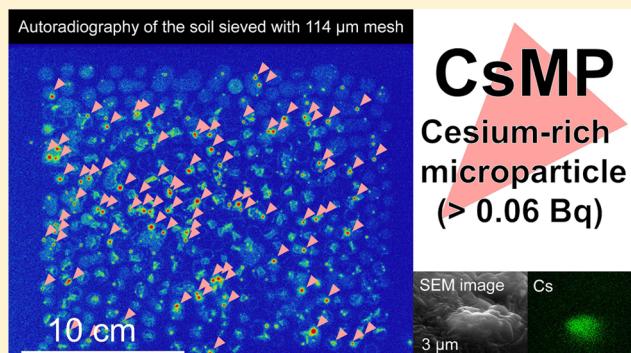
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Supporting Information

ABSTRACT: Highly radioactive cesium-rich microparticles (CsMPs) were released from the Fukushima Daiichi nuclear power plant (FDNPP) to the surrounding environment at an early stage of the nuclear disaster in March of 2011; however, the quantity of released CsMPs remains undetermined. Here, we report a novel method to quantify the number of CsMPs in surface soils at or around Fukushima and the fraction of radioactivity they contribute, which we call “quantification of CsMPs” (QCP) and is based on autoradiography. Here, photostimulated luminescence (PSL) is linearly correlated to the radioactivity of various microparticles, with a regression coefficient of 0.0523 becquerel/PSL/h (Bq/PSL/h). In soil collected from Nagadoro, Fukushima, Japan, CsMPs were detected in soil sieved with a 114 μm mesh. There was no overlap between the radioactivities of CsMPs and clay particles adsorbing Cs. Based on the distribution of radioactivity of CsMPs, the threshold radioactivity of CsMPs in the size fraction of <114 μm was determined to be 0.06 Bq. Based on this method, the number and radioactivity fraction of CsMPs in four surface soils collected from the vicinity of the FDNPP were determined to be 48–318 particles per gram and 8.53–31.8%, respectively. The QCP method is applicable to soils with a total radioactivity as high as $\sim 10^6$ Bq/kg. This novel method is critically important and can be used to quantitatively understand the distribution and migration of the highly radioactive CsMPs in near-surface environments surrounding Fukushima.



INTRODUCTION

The nuclear disaster that occurred after the great east Japan earthquake in March of 2011 caused serious damage to the nuclear reactor unit nos. 1–4 at the Fukushima Daiichi Nuclear Power Plant (FDNPP),¹ which released a total of $\sim 10^{19}$ Bq of radioactivity associated with radionuclides such as ^{133}Xe , ^{132}Te , ^{131}I , and ^{137}Cs .² Of these radionuclides, 10^{15} – 10^{16} Bq of ^{137}Cs was released to the atmosphere and dispersed to the surrounding environment.^{3–5} Even though the fraction of the released Cs was as much as $\sim 7\%$ of the total inventory,³ radioactive Cs is currently responsible for nearly all the

radiation dose in contaminated areas due to its half-lives of 2.065 years and 30.17 years for ^{134}Cs and ^{137}Cs , respectively.⁶ The areas that were subject to Cs contamination span over 15 prefectures including the metropolitan city of Tokyo.^{7–10} The chemical forms of Cs released from the reactor were identified to be soluble hydroxides and chloride¹¹ in addition to possibly

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iodide, which was predicted in accident-code calculations^{12,13} and an experimental study.¹⁴ Kaneyasu et al. reported that the soluble form of Cs reacted with sulfate aerosol and was transported over long distances (>170 km from the FDNPP).¹⁵ The intrinsic forms of soluble Cs and the soluble Cs adsorbed onto the aerosols were then considered to have been deposited on the ground in the large area mentioned above by precipitation, primarily during the periods of March 15th and 16th in the northern Fukushima Prefecture and March 21st–23rd in the Ibaraki, Chiba, Tochigi, and Saitama prefectures and Tokyo.^{8,16}

The deposited Cs subsequently dissolved in the meteoric water, and the hydrated Cs ions quickly adsorbed onto the soil forming minerals and reached adsorption equilibrium within ~10 days.¹⁷ The soluble Cs was efficiently adsorbed to the clay minerals due to its high affinity for edge and frayed-edge sites^{18–21} as well as organic matter. As such, the Cs remains in the surface soil within the top 5 cm of the vertical soil profile in Fukushima due to the tight binding between the Cs⁺, organic matter, and the mineral adsorption sites.⁶ Hereafter, clay minerals adsorbing Cs are referred to as “Cs-clay”. Even though the radioactivities of ¹³⁴Cs and ¹³⁷Cs are as high as ~10⁷ Bq/kg in the soil in Fukushima,⁶ the Cs concentration is only at the level of ~3 × 10⁻¹ ppb due to dilution. Thus, FDNPP-derived Cs is barely detectable using microscopy techniques.

Conversely, several recent studies have characterized and reported a different form of Cs in the soils near the FDNPP. This sparingly soluble form of Cs comes in the form of Cs-rich microparticles (CsMPs), in which the Cs concentration is much higher (~0.55–30 wt % Cs₂O), with the microparticles composed of amorphous SiO₂ and Zn–Fe oxide nanoparticles that are associated with Cs and even minor U and Tc.^{11,22} The nanoscale characteristics and isotopic signatures of these CsMPs provide many clues about the chemical reactions that occurred during the reactor meltdown at the FDNPP and provide a basis for understanding the properties of the debris. Their composition indicates an origin that is a mixture of melted nuclear fuels and structural materials.^{11,22,23} The CsMPs are very small, usually a few micrometers, and have an extremely high level of radioactivity, ~10¹¹ Bq/g, which, in natural water environments or in organisms, can produce high concentrations of hydroxyl radicals, such as •OH, by radiolysis within a thickness of a few hundred micrometers of the water layer adjacent to a CsMP.⁹ Due to the high production rate of hydroxyl radicals and the small size of the CsMPs, internal exposure to a micron-sized point source with extremely high radioactivity is potentially an important health concern.⁹ Therefore, the distribution pattern of the amounts of CsMPs in the contaminated areas needs to be determined on a quantitative basis for better estimation of radiation exposure, especially in locations where residents have returned. Accordingly, in the present study, we focus on the quantification of CsMPs with a radioactivity density of ~10¹¹ Bq/g. This is explicitly different from another type of glassy particle bearing lower concentrations of Cs (undetectable by scanning electron microscopy–energy-dispersive X-ray analysis, SEM-EDX), in which the particle size range was much larger (70–400 μm).²⁴

The CsMPs can be imaged as localized radioactive spots using autoradiography of soil samples.²⁵ In general, the radioactive spots are distributed heterogeneously within the broad distribution of a weak radioactive background that corresponds to the Cs-clay. The highly radioactive spots can be

removed with adjacent soil particles using double-sided tape, and the subsequent γ spectrometry can roughly estimate the radioactivity fraction derived from CsMPs within soil samples with known radioactivity. However, this procedure still involves large uncertainties and produces unreliable results for determining the fraction of radioactivity derived from CsMPs because many Cs-clay particles are also picked up by the double-sided tape, thus introducing large errors and uncertainties in the determination of Cs fraction. More importantly, the Cs-clays are also easily recorded as localized radioactive spots.^{26,27} Even though the adsorbed radioactive Cs is undetectable with electron microscopy at <0.1 wt %, the radioactivity of a large Cs-clay particle and large aggregates of fine clay particles can be high enough to appear as radioactive spots on imaging plates.²⁸ While autoradiography can be used to estimate the radioactivity of individual radioactive particles from Chernobyl²⁹ and Fukushima,^{30,31} its application to the quantification of CsMPs has never been explored. As reported in previous studies, individual CsMPs can be separated and characterized in detail using electron microscopy techniques and γ spectrometry;¹¹ however, it is impossible to separate and characterize all individual CsMPs in a given sample because the number of radioactive spots is too large to count and the efficiency of successful separation is not 100%. Therefore, the development of an efficient and simplified method to quantify the number and radioactivity of CsMPs is important. Here, we describe a novel application and significant expansion of a classic measurement technique (autoradiography) to quantify the fraction of radioactivity contributed by the CsMPs in contaminated soils. We then apply the new method to four surface soil samples in Fukushima to confirm the validity of the results.

SAMPLING AND METHODS

Sample Description. Figure 1 shows the Cs radioactivity distribution around the FDNPP³² and the sample locations for

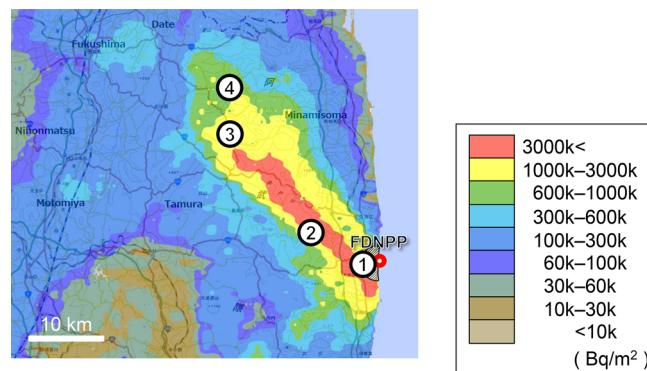


Figure 1. Distribution of Cs¹³⁴⁺¹³⁷ radioactivity and sample locations. The map was modified with permission from Extension Site of Distribution Map of Radiation Dose, etc./GSI Maps.³² The number represents the sampling points: no. 1, Ottozawa; no. 2, Omaru; no. 3, Nagadoro; and no. 4, Fukaya.

this study. During sampling campaigns in 2011 and 2012, surface soil samples in fields devoid of vegetation were collected from four locations in the Fukushima Prefecture: Ottozawa (no. 1), Omaru (no. 2), Nagadoro (no. 3), and Fukaya (no. 4), all being within a distance of 4.42–40.3 km from the FDNPP (Figure 1 and Table 1). These soil samples are mainly composed of typical soil minerals such as quartz, feldspars,

Table 1. Information Concerning the Samples Used in the Present Study to Which the QCP Method Was Applied

sampling point	location (sample number in KU sample archive)	latitude-longitude	distance from FDNPP (km)	soil type	sampling date
no. 1	Ottozawa (KU171)	37°25'10.045" N, 141°00'25.387" E	4.42	surface soil of field	March 16, 2012
no. 2	Omaru (KU1)	37°28'00.490" N, 140°55'47.514" E	13.0	surface soil of field	December 20, 2012
no. 3	Nagadoro (KU133)	37°36'44.913" N, 140°44'56.777" E	32.3	surface soil of paddy	March 15, 2012
no. 4	Fukaya (KU39)	37°42'43.699" N, 140°43'36.499" E	40.3	surface soil of field	October 25, 2012

micas, and clay minerals. The details of soil characterization can be found in previous reports.^{6,11,17,33} The amount of adsorption sites in clay minerals is large compared to that of soluble radioactive Cs. The calibration and determination of the accuracy analysis of the autoradiography protocol were conducted using several radioactive particles: one Cs-clay, three gravels, and five CsMPs. The characteristics of these particles, including radioactivity, have already been reported in our previous study,^{22,23} and this is also summarized in Table 2.

Table 2. List of the Samples Used as Point Sources for the Calibration of the Radioactivity Measurement Using Autoradiography^a

sample name	radioactivity (Bq)	size (maximum diameter)	type
NGD-clay1	0.228	0.50 mm × 60 μm height	Cs-clay
OTZ8	3.64	2.6 μm	CsMP
OTZ9	4.89	2.0 μm	CsMP
KOI2	17.8	4.4 μm	CsMP
OTZ7	72.1	5.8 μm	CsMP
G2	135	0.97 mm	gravel
G5	240	2.0 mm	gravel
G3	259	2.7 mm	gravel
OTZ3-before ^b	418	17 μm	CsMP
OTZ3-after ^b	334	17 μm minus the size of the pit by the SIMS analysis	CsMP

^aNGD-clay1 and OTZ7 were used to draw the PSL intensity–t diagram, and all samples were used to deduce the PSL/h-D graph. Detailed microscopic characteristics of some grains have been reported elsewhere. ^bFrom Imoto and Ochiai et al.²² OTZ3-before and -after stand for the same CsMP before and after the SIMS analysis.

Safety. ¹³⁴Cs and ¹³⁷Cs are radioactive β-emitters with γ-emitting daughter isotopes. Radioisotopes should be handled by suitably qualified and experienced personnel in a properly equipped laboratory, and any work should follow appropriate risk assessment. The possession and use of radioactive materials is subject to statutory controls.

γ Spectrometry. The ¹³⁴Cs and ¹³⁷Cs radioactivities of the individual CsMPs were determined using γ spectrometry. Here, a point-source specimen, which was described in a previous study,¹¹ was used for standardization. Briefly, the radioactivity of a microparticle with a size of ~400 μm obtained from the surface soil in Fukushima was precisely determined at the radioisotope center in Tsukuba University, Japan, and utilized as a standard point specimen for ¹³⁴Cs and ¹³⁷Cs. The radioactivity of the point source standard was 23.9 Bq for ¹³⁴Cs and 94.6 Bq for ¹³⁷Cs as of September 29, 2015. The radioactivity measurements for the other particles and Cs clay were performed using SEIKO E&G germanium semiconductor detectors GMX23, GMX30, and GMX40 at Kyushu University.

city's Radioisotope Center. The acquisition time ranged from 1 to 162 h to ensure that the net area of both the ¹³⁴Cs and ¹³⁷Cs peaks exceeded 1000 counts.

Quantification of CsMPs Method. A method to quantify the CsMPs, hereafter called the QCP method, was developed using the photostimulated luminescence (PSL) of an imaging plate (IP, FUJIFILM BAS-SR 2025) with an IP reader (GE, Typhoon FLA 5100). The pixel size was set to 50 μm. The PSL intensity is proportional to the radioactivity and exposure time.³⁴ First, the relationship between the PSL intensity and the exposure time (herein PSL/h) was determined using two radioactive particles; a micron-sized clay particle that had adsorbed radioactive Cs (NGD-clay1) and a CsMP (OTZ7). These samples provided different total γ radiation owing to their ¹³⁴Cs and ¹³⁷Cs content (0.228 ± 0.003 and 72.1 ± 0.7 Bq for NGD-clay1 and OTZ7, respectively). The exposure time varied from 15 min to 73 h. The radioactivity of the standard microparticles was determined using γ spectrometry.

Second, the relationship between the PSL intensity per hour and the radioactivity (in becquerel) was established by comparing six microparticles, one Cs-clay, and three large millimeter-sized gravels, whose radioactivities were previously determined to be different from each other using γ spectrometry (Table 2). The exposure time for the PSL measurement was set such that the total PSL count became less than the upper limit, and a linear relation between the PSL intensity per hour and the radioactivity was retained.

Distinguishing between Cs-Clay and CsMPs by Setting a Threshold Radioactivity. A soil sample collected from Nagadoro was used to determine the threshold radioactivity for the CsMPs. First, 2.95 g of the soil sample was sieved with 597 and 114 μm meshes in a fume hood and separated into three particle size fractions: >597 μm (large; 0.86 g), 114–597 μm (medium; 1.2 g), and <114 μm (small; 0.88 g). For the large-size fraction, radioactive particles were picked from the sample after autoradiography to isolate them. For the medium- and small-size fractions, the radioactive particles were separated according to the procedure for separating CsMPs from soil samples reported in our previous study.¹¹ Briefly, the powder samples were dispersed on grid paper and covered with a plastic sheet, and an IP was then placed on the samples for 1 hour. After the positions of intensely radioactive spots were identified, droplets of pure water were added to these positions and then drawn using a pipet to produce suspensions with small amounts of soil particles by dilution with pure water. This procedure was repeated until the suspension did not contain a significant amount of soil particles. Subsequently, the isolated radioactive particles were characterized in detail to identify the Cs-clay, CsMPs, or other aggregates by means of the following procedures. (i) The Cs-clay was identified based on the platy

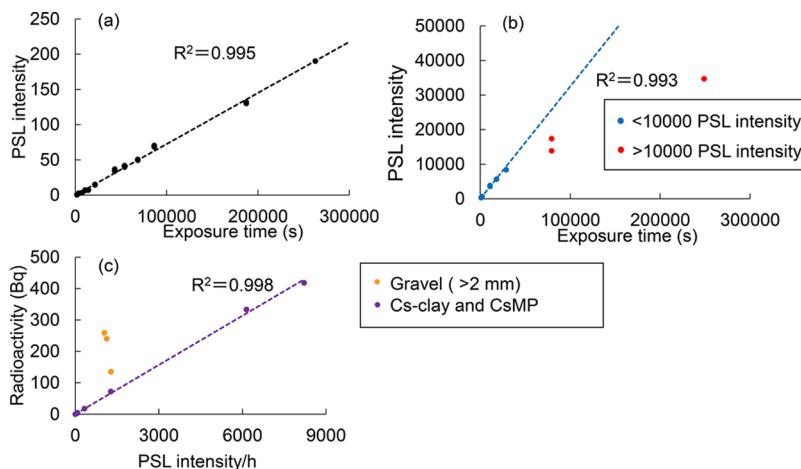


Figure 2. Correlation between parameters in the autoradiography analysis. (a) Correlation between the PSL intensity and the exposure time, which was measured using a Cs-clay (NGD-clay1). The regression line is shown as a dotted line. (b) Correlation between the PSL intensity and the exposure time, which was measured using a CsMP (OTZ7). The dotted line represents the regression line determined using the plots with <10 000 PSL counts. (c) Correlation between the radioactivity and the PSL/h, based on measurements of one Cs-clay, six Cs MPs, and three gravels. The dotted line indicates the regression line created using the data for the Cs-clays and the CsMPs.

shape of micaceous minerals in the stereomicroscope image. (ii) To separate CsMPs that adhere to visible clay particles, the visible and radioactive clay particles were subjected to an ultrasonic treatment in ultrapure water for 20 min, and then autoradiography was performed. It is noted that Cs in both clays and CsMPs scarcely dissociate in water during the sonication because it tightly binds to clay minerals³³ or occurs within SiO₂ glass matrix of CsMP. A visible clay particle was determined to be a Cs-clay if the radioactive spot remained at the position of the clay particle; however, it was considered to be a CsMP if the radioactive spot was dissociated from the clay particle. (iii) The radioactive particle was crushed using a blade, and autoradiography was then performed. The platy particle and aggregate were judged to be a Cs-clay and aggregate when the radioactive spot was separated along with the crushed fragments,²⁸ whereas it was determined to be a CsMP if it remained as a single spot in the autoradiography image. (iv) The radioactive particles were found and characterized by scanning electron microscopy (Shimadzu, SS550 and Hitachi, SU6600) using an energy dispersive X-ray analysis (EDAX Genesis). Information on the shape and composition of the clay and aggregate were obtained. For the particular case, the CsMPs were also searched using SEM-EDX, although it was difficult to detect and characterize all of the CsMPs. The above four procedures were employed as appropriate for each sample.

After differences between CsMPs and Cs-clay were distinguished, a histogram of the particles was produced as a function of the radioactivity for the three size fractions (large, medium, and small) as described above. The radioactivity of each radioactive particle was calculated based on the QCP method described in the previous section. The threshold radioactivity was determined for the identification of the CsMPs using the histogram.

Application of the QCP Method to Surface Soils. Approximately 0.5 g of four soil samples from Ottowaza, Otaru, Nagadoro, and Fukaya, listed in Table 1, were subjected to ultrasonic treatment for 30 min and sieved with 597 and 114 µm meshes. Next, the QCP method was applied. In the autoradiograph image of the small fraction, all spots were determined to be either Cs-clay, aggregate, or CsMP using the radioactivity threshold. The radioactivity fraction (RF) of the

CsMPs (i.e., the amount of radioactivity that can be attributed to the CsMPs) can be calculated using the following equation:

$$RF = \frac{D_{CsMP}}{D_{all}} \times 100(\%)$$

Here, D_{CsMP} is the sum of the CsMP radioactivity (in becquerel) in the soil samples determined by the QCP method, and D_{all} represents the total radioactivity (in becquerel) of the soil sample determined by γ -ray spectrometry. The number of CsMPs in the soil samples was quantified by counting the radioactive spots greater than the radioactive threshold.

RESULTS

Correlation of the PSL Intensity and Radioactivity.

Figure 2a shows the linear correlation between the PSL intensity and the exposure time for the measurement of the Cs-clay (NGD-clay1) with low radioactivity, 0.228 ± 0.003 Bq. A linear correlation between the PSL intensity and the exposure time was also obtained for the measurement of the CsMP (OTZ7) when the PSL intensity was less than 10 000 (Figure 2b); however, the PSL plots deviated from the regression line of the correlation when the PSL intensity exceeded 10 000, which is approximately consistent with the maximum counts of the dynamic range of the imaging plate as certified by the manufacturer. Therefore, the upper limit of the PSL count was determined to be 10 000. For the correlation line between the PSL intensity and the exposure time, the standard deviation was calculated to be $\sigma = 3.94$, which results in a limit of quantification for the PSL measurement of $10\sigma = 39.4$. In Figure 2c, the correlation between the PSL intensity per hour (PSL/h) and radioactivity (becquerel) is plotted for Cs-clays and CsMPs that have a small particle size (approximately the micron scale) and for gravel samples with a particle size of $>\sim 2$ mm. The gravel sample data deviates from the regression line, which is attributed to the large geometric variation in the distance between the imaging plate and the samples. Based on the regression line excluding all the data for the gravel samples, a correlation of $0.0523 (\pm 0.0008)$ Bq for 1.00 PSL/h ($R^2 = 0.998$) was obtained; this was subsequently used for the

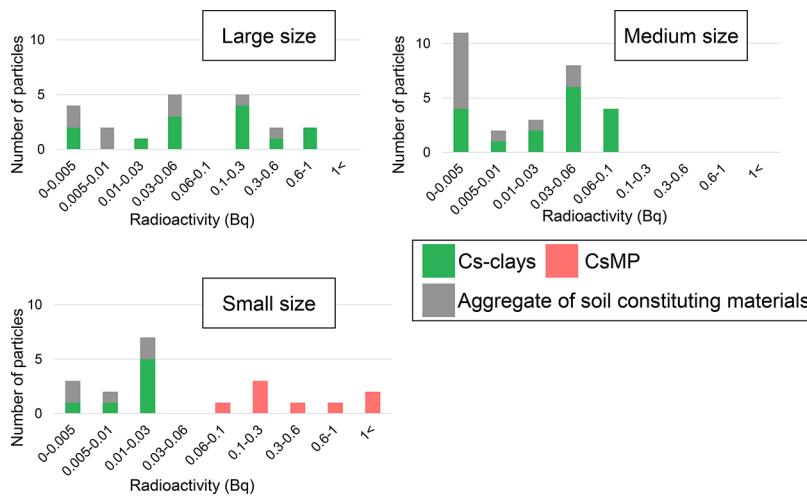


Figure 3. Histogram of the CsMP and Cs-clay distributions for three size fractions. Large, medium, and small size fractions are >597 , $114\text{--}597$, and $<114\text{ }\mu\text{m}$, respectively.

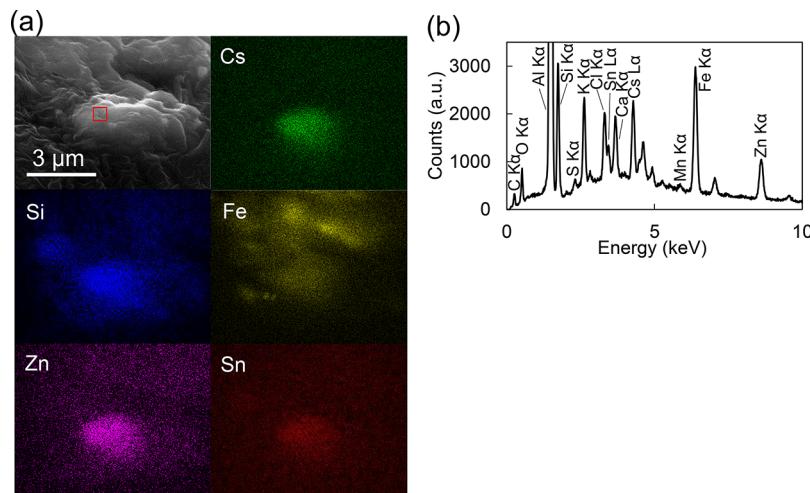


Figure 4. Scanning electron microscopy of the CsMP isolated from the Nagadoro soil sample. (a) Secondary electron (SE) image of a CsMP associated with the energy dispersive X-ray spectrum (EDX) elemental maps of the major constituents. (b) EDX spectrum of the area indicated by the red square in the SE image in panel a.

conversion of the PSL intensity to radioactivity in the following experiments. Note that the conversion factor is valid only when conditions are the same as the present experiment (i.e., FUJIFILM BAS-SR 2025 for the imaging plate, GE Typhoon FLA 5100 for the IP reader, and $50\text{ }\mu\text{m}$ for the pixel size). When different experimental conditions are used, the parameters such as the pixel size, the exposure time, the maximum PSL count, and PSL intensity–radioactivity correlation need to be recalibrated.

Threshold Radioactivity for Distinguishing Cs-Clay and CsMPs. A total of 69 radioactive particles were isolated from the soil sample collected from Nagadoro and further investigated for their size, radioactivity, and whether they were CsMPs. We distinguished whether the isolated particles were CsMPs, Cs-clays, or aggregates based on the procedure described in the *Sampling and Methods* section. Based on the identification of all 69 particles, it was concluded that no CsMPs are present within the medium ($114\text{--}597\text{ }\mu\text{m}$) and large ($>597\text{ }\mu\text{m}$) size fractions, whereas CsMPs were identified in the small fraction ($<114\text{ }\mu\text{m}$) with a radioactivity of $>0.06\text{ Bq}$ (Figure 3). Therefore, the criteria to identify CsMPs is that the

particle is $<114\text{ }\mu\text{m}$ with a radioactivity $>0.06\text{ Bq}$ for the present experimental setting. A CsMP was isolated from the fraction satisfying these criteria and characterized using a SEM–EDX (Figure 4). A secondary electron image shows an aggregate of the CsMP and the other soil particles attached on the double-sided carbon tape (Figure 4a). The composition of the CsMP is similar to the ones previously reported,¹¹ its main constituents are oxides of Si, Fe, Zn, Sn, and Cs (Figure 4a,b), although the EDX spectrum has an Al signal derived from Al substrate and some other elements including Ca, K, and Cl as well as small interfering peaks from the constituents of clay and Fe oxide background. The ^{134}Cs and ^{137}Cs radioactivities of the CsMP were determined to be 1.90 and 1.85 Bq, respectively (Table 3). The radioactivity was decay-corrected to March 12, 2011, 15:36 Japan standard time (JST), when the first hydrogen explosion occurred. The ^{134}Cs -to- ^{137}Cs radioactivity ratio was 1.02 ± 0.05 , confirming that the CsMP was released from the FDNPP.³⁵

Application of the QCP Method to Surface-Soil Samples.

Our newly developed QCP method was applied to four surface soil samples (Table 1). The autoradiography

Table 3. Radioactivity of a CsMP Discovered in the Nagadoro Soil Sample^a

particle size (μm)	radioactivity (Bq) ^b		
	^{134}Cs	^{137}Cs	$^{134}\text{Cs}/^{137}\text{Cs}$
1.18	1.90 (± 0.10)	1.85 (± 0.01)	1.02 (± 0.05)

^aThe ^{134}Cs and ^{137}Cs radioactivities of an isolated individual CsMP measured using γ -spectrometry. ^bThe radioactivity was decay-corrected to March 12, 2011, 15:36 JST.

images for the three size fractions were collected using the QCP method for all four samples, and the images for a Nagadoro sample are shown in Figure 5 as representative data. The autoradiography images of the other samples are compiled in Figures S1–S3. Note that sieving with the 597 and 114 μm meshes for the QCP method efficiently excluded the Cs-clay from the CsMP group. The pink arrows in the small fraction indicate the radioactive spots that satisfy the criteria described above and are identified as CsMPs (Figure 5). Each spot was manually counted and is indicated by an arrow. The sum of the radioactivity of all the spots was calculated to be 16.1 ± 0.1 Bq. The total radioactivity derived from the CsMP identified in the sample was used to calculate the RF of the CsMP. As a result, the RF of CsMPs in Nagadoro soil (as per Figure 5) was determined to be $19.8 \pm 0.3\%$. The total number of radioactive spots identified as CsMPs was determined to be 154 ± 4 CsMPs per gram in the Nagadoro soil (see Table 3). The RF and the number of CsMPs in all four soil samples (Figures 5 and S1–S3) were determined to be 8.53–31.8% and 48–318 CsMPs/g, respectively, indicating that the RF and the number varies greatly depending on the location (Table 4). In addition, the effect of sample volume on the QCP method was investigated using the soil with the lowest radioactivity, ~6.8 Bq/g, in which a limited number of CsMPs were present. The results of QCP method are summarized in Table S2. The radioactivity fractions of CsMPs were calculated to be 23% and 27% for 7 and 0.5 g, respectively. The results obtained here are in good agreement with each other within the reasonable error range. Thus, the 0.5 g sample mass is a reasonably optimized mass for the application of the QCP method to the samples with total radioactivity as low as ~7 Bq/g. Increasing the mass of the soils being analyzed would permit extension of the applicability of the technique to soils with radioactivities much lower than ~7 Bq/g.

DISCUSSION

Validity of the Exclusion Procedure for Clay Particles Adsorbing Radioactive Cs.

As mentioned above, the Cs concentration on the Cs-clay fraction is considered to be low; however, the radioactivity density of the clay has not been well evaluated. In Koarashi et al.,³⁶ assuming that all Cs is associated with clay fraction, the ~60 kBq/kg maximum radioactivity corresponds to only 3–12 Bq/g clay based on their volume percentage of 5–20%. In contrast, another study³⁷ reported that clay particles with a particle size of ~50 μm can sorb Cs as much as 10^{-2} Bq.³⁷ Thus, there is a possibility that the highest Cs radioactivity on clay could exceed the threshold value. To confirm and determine the applicability of the QCP method, we have isolated radioactive particles from the Ottozawa soil, and the most-radioactive clay particle was isolated for further SEM observation (Figure S4). The volume of this particle was approximated based on the two SEM images tilted at 0° and 45°. The composition was qualitatively analyzed by SEM–EDX and was similar to that of weathered biotite. The radioactivity was determined to be 0.514 ± 0.013 Bq using γ spectroscopy. The radioactivity density was calculated to be $\sim 1.3 \times 10^4$ Bq/g assuming the density of clay as 2.8 g/cm^3 . Here we assume that the radioactivity density can be applied to micron-sized clay particles in the present study because Mukai et al. reported that radioactive Cs is uniformly distributed on weathered biotite particles as small as ~100 μm that were collected from Fukushima.²⁸ Based on this calculated value, the validity of the method distinguishing between CsMPs and Cs-clays is discussed below.

The present analytical method determined the criteria for finding CsMPs: smaller than 114 μm and a threshold radioactivity of 0.06 Bq, which was the lower limit of the range of radioactivity for CsMPs in the frequency histogram. Here, assuming a Cs-clay particle that has the maximum radioactivity and is 114 μm in size, the radioactivity of a Cs-clay (D) can be calculated as follows:

$$D = D_d \times d \times V$$

where D_d is the highest radioactive density of Cs-clay determined above ($\sim 1.3 \times 10^4$ Bq/g), d is the mass density of Cs-clay (in which the typical density of clay minerals, $\sim 2.8 \text{ g/cm}^3$, is adopted),³⁸ and V is the largest volume of Cs-clay that can pass through a 114 μm mesh. To calculate the volume, a hexagonal polyhedron with a longest diameter of 114 μm (the length of one side is 57 μm) with a height of 114 μm was considered, where the volume was calculated to be $\sim 1 \times 10^{-6} \text{ cm}^3$. This volume gave a maximum possible radioactivity of an

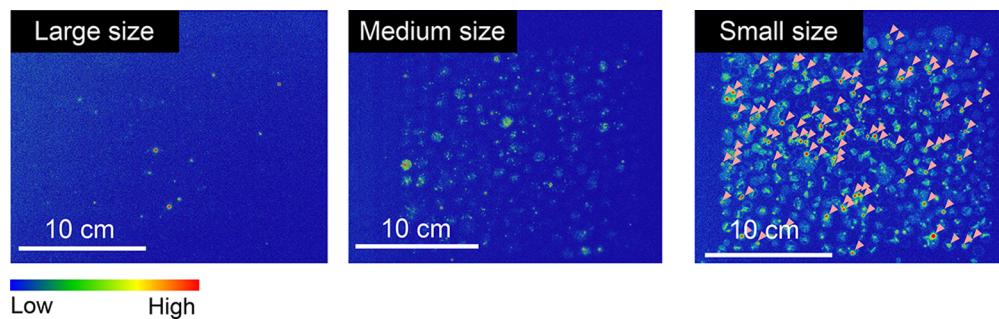


Figure 5. Representative result of the QCP method for the Nagadoro soil sample. Color-scaled autoradiography images of the large, medium, and small size fractions indicating the sizes of the particles (>597, 114–597, and <114 μm , respectively). The pink arrows indicate the CsMP spots, where the radioactivity was determined to be higher than the threshold value (0.06 Bq). The exposure time was 5 days.

Table 4. Results of the QCP Method Applied to Soil Samples from Ottozawa, Omaru, Nagadoro, and Fukaya

area	mass (g)	number of CsMPs (particles per gram) ^a	total radioactivity (Bq/g) ^b	total radioactivity of CsMPs (Bq/g) ^a	radioactivity fraction (RF) of CsMPs (%)
Ottozawa	0.453	318 (± 18)	1016 (± 10)	86.8 (± 1.1)	8.53 (± 0.14)
Omaru	0.532	165 (± 5)	236 (± 1)	44.9 (± 0.3)	19.0 (± 0.1)
Nagadoro	0.539	154 (± 4)	151 (± 2)	29.9 (± 0.3)	19.8 (± 0.3)
Fukaya	0.554	48.2 (± 3.7)	46.6 (± 0.2)	14.8 (± 0.2)	31.8 (± 0.5)

^aThe number and total radioactivity of the CsMPs were measured using the QCP method. ^bThe total radioactivity of the soil samples was analyzed via γ -spectrometry.

individual Cs-clay of $D = \sim 0.036$ Bq, which is lower than the threshold radioactivity of a CsMP, 0.06 Bq. This calculation clearly indicates the validity of this procedure for distinguishing between CsMPs and Cs-clay in the present experiment. Thus, it is assured that the QCP method is applicable to the soils contaminated with radioactive Cs, of which the bulk radioactivity is as high as that of the Ottozawa soil, 10^3 Bq/g = 10^6 Bq/kg (Table 4). However, it is plausible that the soil samples that have greater total radioactivity than that of Ottozawa contain clay particles with a radioactivity density higher than the threshold value of CsMP. Thus, the QCP method can be applicable to the soil with total Cs radioactivity of $< 10^6$ Bq/kg, which actually covers most of the area of the Fukushima prefecture. It cannot be applicable only to soil samples with total Cs radioactivity concentration (hot spot) near the FDNPP. Thus, the merit of QCP method still stands, and it is important for the quantification of CsMPs in ambient environments.

Furthermore, the procedure of the present QCP method can be easily optimized for the application to the soils with higher total radioactivity than 10^6 Bq/kg. Using the sieving mesh with smaller pore size than 114 μm (e.g., a 10 μm sieve) would exclude the presence of Cs-clay with higher radioactivity than the threshold value of CsMP. Thus, the modified QCP can be used to quantify the number of CsMPs in sample regardless of the total radioactivity.

Evaluation of the Particle Diameter of CsMPs that Correspond to the Threshold Radioactivity. The threshold radioactivity determined in the present study is the minimum radioactivity of CsMPs that can be separated from the soil; however, there is still ambiguity if the threshold corresponds to a reasonable size of the CsMPs. Therefore, the diameter of a particle having the threshold radioactivity was estimated as follows.

Assuming the shape of CsMP as sphere, the particle diameter of a CsMP having the threshold radioactivity (Φ_{th}) can be expressed as:

$$\Phi_{\text{th}} = 2 \left(\frac{3V}{4\pi} \right)^{1/3}$$

$$V = \frac{D_{\text{th}}}{D_{\text{d,max}} d \times 10^{12}}$$

where D_{th} is the threshold radioactivity of a CsMP (0.06 Bq); $D_{\text{d,max}}$ is the maximum radioactivity density of the CsMPs that have been characterized so far ($7.77 \pm 0.09 \times 10^{11}$ Bq/g); d is the density of the CsMPs, which was set to 2.65 g/cm³ based on the similar chemical properties of SiO₂ glass;³⁹ V is the volume of the CsMP; and Φ_{th} is the diameter of the CsMP with the threshold radioactivity. The calculation resulted in 0.4 μm for Φ_{th} .

Figure 6 illustrates the distribution of the radioactivity of CsMPs that have < 5 Bq as a function of the particle size, as

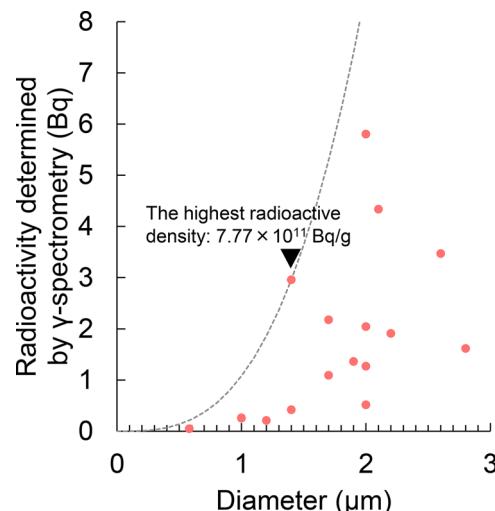


Figure 6. Radioactivity of CsMPs as a function of their size. A summary of the radioactivities of CsMPs as a function of their particle size for the particles that have been discovered to date.^{9,11,22,23} The dotted curve indicates the ideal relation between the radioactivity and the size for CsMPs having the maximum radioactivity density, assuming a homogeneous radioactivity density within the particle.

reported by previous studies.^{9,11,22} As can be seen in the diagram, the smallest CsMP reported so far is 0.58 μm in diameter,⁹ which is larger than Φ_{th} . This result may indicate low probability of finding CsMPs $< \sim 0.5 \mu\text{m}$ in the environment. The histogram of the particle's radioactivity in Figure 3 reveals that the distribution of CsMPs is clearly separated from that of Cs-clay particles, and there are no CsMPs with radioactivity less than 0.06 Bq. Therefore, it is likely that CsMPs released into the environment are generally larger than $\sim 0.5 \mu\text{m}$. This hypothesis is partly supported by a previous report by the Nuclear Energy Agency,⁴⁰ which theoretically calculated the aggregation of particles smaller than 0.1 μm during the molten core-concrete interactions in a severe accident. A previous experimental study of molten core-concrete interactions also revealed the formation of microparticles as aggregates of 0.1 μm sized clusters.⁴¹ Indeed, previous transmission electron microscopy images revealed evidence of aggregation within a single particle.^{22,23} It is noted that in the presented method, aggregates that cannot be separated by the sonication process would be considered a single particle.

Implications of the QCP Method for Evaluating CsMPs in the Biosphere. The QCP method developed in the present study is a simple but powerful method of quantifying the amount of radioactivity and number of CsMPs in soil samples, even though the threshold values need to be reset depending

on the analytical protocol, such as the IP and the IP reader. The RF value of the CsMPs was determined to be as high as 32% in one soil sample, proving the significant contribution of CsMPs to the total radioactivity in the FDNPP-affected environment. The number of soil samples analyzed in the present study was too limited to draw a detailed quantitative picture of the CsMP distribution. Constructing a map of the CsMP distribution by analyzing a larger set of samples would provide quantitative insight into understanding the timing of the formation, release, and movement of CsMPs from the FDNPP during the initial stage of the nuclear disaster. Due to their highly localized radioactivity, the quantity of CsMPs has been of great concern in residential areas in the Fukushima Prefecture. The RF values determined using the QCP method can be used as a representative indicator to convert the total Cs radioactivity to the numbers of CsMPs in each residential area near the data points. In addition, the contribution of CsMPs to the migration of radioactive Cs in the environment after the nuclear disaster has largely been overlooked as are their potential role in human metabolism. Although it is difficult to directly link the RF value and the number of CsMPs with the health impacts, those data are essential for the subsequent calculation of the radiation exposure. The number of the CsMPs is a measure of potential resuspension to air. Some fraction of the resuspended CsMPs can potentially be deposited deep in the respiratory system.^{9,42}

It is important to predict the redistribution of the radiation dose over decadal time periods considering CsMPs. Indeed, Cs migration in the form of CsMPs is another important route for Cs transport in the environment, which can be tracked using the QCP method. Simple distribution coefficient, K_d or surface-complexation models will not be useful for the description of the migration of this form of Cs in the environment, in contrast to the ionic forms of Cs.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: [10.1021/acs.est.7b06693](https://doi.org/10.1021/acs.est.7b06693).

Tables showing radioactivity of the individual particles and results of the QCP method. Figures showing the results of the QCP method and secondary-electron images and elemental maps. ([PDF](#))

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Author Contributions

S.U. conceived the idea and designed all experiments. R.I. conducted QCP analysis. R.I. and S.U. wrote the manuscript. A.O., M.S., T.K., and G.F. performed the separation of CsMPs and SEM analysis. T.O. provided navigation during field research in Fukushima. S.Y. performed γ spectroscopy at Tsukuba University. W.R.B., G.T.W.L., B.G., and R.C.E. participated in the discussion and interpretation of the results.

Notes

The authors declare no competing financial interest.

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